WORKSHOP 3: Name: _____
Nomenclature

Section _____

In each box, write the correct formula for the ionic compound which contains the positive ion (cation) listed at the top of the column and the negative ion (anion) listed on the left.

	Li^+	**Mn^{2+}**	**Fe^{3+}**	**Zn^{2+}**	**NH_4^+**
I^-	(1)				
S^{2-}		(2)			
N^{3-}			(3)		
OH^-				(4)	
SO_4^{2-}					(5)
NO_3^-				(6)	
PO_3^{3-}			(7)		
CO_3^{2-}		(8)			

Name the compounds in the boxes above with the bracketed numbers (1) , (2) etc

1. _____ 2. _____

3. _____ 4. _____

5. _____ 6. _____

7. _____ 8. _____

Name the Compounds:

CuI	_____
$NaNO_3$	_____
$Sr(OH)_2$	_____
CoO	_____
NH_4F	_____
$LiBrO_4$	_____
SO_3	_____
SnS_2	_____
$SiBr_6$	_____
$FePO_4$	_____
AgCN	_____
$Be(HCO_3)_2$	_____
HI(aq)	_____
$Ca(ClO)_2$	_____
Hg_2CO_3	_____
HNO_2	_____
N_2O_4	_____
$PbSO_4$	_____
AlH_3	_____
$HC_2H_3O_2$	_____
$Sn(ClO_3)_4$	_____
GaP	_____
XeF_4	_____

Write Formulas:

cadmium (II) carbonate	_____
ozone	_____
carbonic acid	_____
lead (II) hydroxide	_____
titanium (II) nitrate	_____
ammonia	_____
barium arsenide	_____
aluminum nitride	_____
ammonium oxide	_____
hydrofluoric acid	_____
chromic acid	_____
lithium phosphate	_____
hydrogen selenide	_____
tin (IV) periodate	_____
methane	_____
carbon tetrachloride	_____
manganese (III) hydroxide	_____
hypoiodous acid	_____
perchloric acid	_____
cupric acetate	_____
gold (III) sulfate	_____
hydrosulfuric acid	_____
nitrogen monoxide	_____